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Chapter 9: Section 2: Ideal Stoichiometric Calculations
9-2 Ideal Stoichiometric Calculations Ideal Stoichiometry - All reactants are converted into products I. A Common Method for Solving All Stoichiometry Problems A. Mass-Mass Problems 1. Start with a known mass of reactant or product, find an unknown mass of another reactant or product 2. All other stoichiometry problems are derivations ...

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Stoichiometry. SECTION 2. PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. The following equation represents a laboratory preparation for oxygen gas: 2KClO3(s) (2KCl(s) + 3O2(g) How many moles of O2 form if 3.0 mol of KClO3 are totally consumed? 2. Given the following equation: H2(g) + F2(g) (2HF(g

CHAPTER 9 REVIEW
93.9% if the percentage yield for the reaction represented by the following equation is calculated to be 75.3%, what mass of Al is expected from the reaction of 52.5g of Al_2O_3 $2Al_2O_3(l) \rightarrow 4Al(s)+3O_2(g)$

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CHAPTER 9 REVIEW Stoichiometry SECTION 3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N 2 are mixed with 12.0 mol of H

mc06se cFMsr i-vi
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